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**Proton Exchange Membrane Materials Based on Sulfonated
Poly(phthalazinones)**

10/529020**JC17 Rec'd PCT/PTO 24 MAR 2005****5 BACKGROUND OF THE INVENTION**

In recent years, polymer electrolyte fuel cells have been identified as promising power sources for vehicular transportation and other applications requiring clean, quiet and efficient portable power. As a vital part of the fuel cell, proton exchange membranes (PEM)s have become a rapidly growing area of research. Until now, PEMs have been mainly limited to perfluorinated ionomer membranes such as Nafion™ developed by DuPont and similar membranes commercialized by Dow and Asahi. In spite of their outstanding properties such as excellent proton conductivity and oxidative resistance, which are essential for fuel cell application, the perfluorinated ionomer membranes are very expensive (US\$800-2000/m²) and suffer from other serious drawbacks of high methanol permeation and dehydration. In order to develop alternatives to Nafion™, which would be less expensive and free from other disadvantages of perfluorinated ionomer membranes, several attempts have been recently made including synthesis of new polymer electrolytes¹⁻⁵, chemical modification of available high performance polymers⁶⁻¹⁰, and blend membranes¹¹⁻¹⁴. Among these studies, introducing sulfonic acid groups into the main chains of high performance polymers by sulfonation reaction is an important and widely used method for imparting polymers with proton exchange capability. Poly(aryl ether ketone)s, poly(phenylene oxide), poly(phenylene sulfide), poly(aryl ether sulfone), and polybenzimidazole are among those that have been studied. The sulfonating agents include concentrated sulfuric acid, chlorosulfonic acid, pure or complex sulfur trioxide, and acetyl sulfate. As an example, the sulfonation reaction of Victrex™ PEEK and the conductivity of sulfonated PEEK have been studied extensively^{7, 11, 15-17}. The DS could be controlled by reaction time and temperature in concentrated sulfuric acid or oleum. For sulfonated PEEK with a relatively low DS of 0.65, its conductivity reaches 0.04 S cm⁻¹ at 100°C/100% RH, higher than that of Nafion-117 measured under the same conditions.

Poly(phthalazinone ether ketone) (PPEK) is one of a new class of poly(aryl ether ketone)s under consideration for commercialization. PPEK has a very high glass transition temperature of 263°C, excellent high-temperature stability, and many other good properties. The sulfonation reactions of its copolymer, poly(phthalazinone ether sulfone ketone) (PPESK), and the nanofiltration and ultrafiltration membrane properties have also been studied¹⁸⁻²³.

SUMMARY OF THE INVENTION

According to the invention, the sulfonation reaction of several poly(phthalazinones) have been investigated, and suitable PEMs acquired from the sulfonated products.

According to one aspect of the invention, a series of sulfonated poly(phthalazinone)ether ketones (PPEKs), (SPPEKs) with different DSs were prepared via modification of PPEK with the mixture of concentrated sulfuric acid and fuming sulfuric acid as the sulfonating agent. The structure and some properties of SPPEKs and resulting membrane films were characterized, including: sulfonation position, proton conductivity, equilibrium water uptake, swelling ratio and thermostability as a function of DS.

Similarly, a series of poly(phthalazinone) ether sulfone ketones (PPESKs), and poly(phthalazinone) ether sulfones (PPESs) were prepared and tested.

BRIEF DESCRIPTION OF THE DRAWING

Figure 1 is a schematic illustration of the synthetic pathway for SPPEK;
Figure 2 is a ¹H-NMR spectrum of PPEK in CDCl₃;
Figure 3 is a ¹H-NMR spectrum of SPPEK in DMSO-d₆;
Figure 4 is a ¹³C-NMR spectrum (hydrogen decoupled) of PPEK in CDCl₃;
Figure 5 is a FT-IR spectra of PPEK and SPPEK;
Figure 6 is a graph illustrating TGA traces of PPEK and SPPEKs;

Figure 7 is a graph illustrating the Conductivity of SPPEK;
Figure 8 is a graph of Conductivity versus temperature for SPPEK and SPPEK; and
Figure 9 is a graph of Conductivity of SPPEKs of various DS.

5 DETAILED DESCRIPTION OF THE INVENTION

EXPERIMENTAL

Materials

PPEK was synthesized according to the procedure reported previously^{18,19}. All other
10 chemicals obtained commercially were reagent grade and used as received.

Sulfonation Reaction

In a typical small-scale experiment, 0.5 g PPEK powder was added to a 10 mL
mixture of 95-98% concentrate sulfuric acid and 27-33% fuming sulfuric acid under
an argon atmosphere and the mixture was magnetically stirred at room temperature
15 (23°C) for a certain reaction time. The reported reaction time is the total time for
polymer dissolution and reaction. After a determined reaction time, the reaction
medium was poured onto crushed ice and the resulting precipitate was recovered by
filtration, washed with deionized water until the pH value was ~6-7.

For scaled-up reactions of 30-35 g, an ice bath was needed initially to cool the
20 heat released during dissolution. An amount of 30 g PPEK powder was added to a
mixture of 240 mL 95-98% concentrated sulfuric acid and 360 mL 27-33% fuming
sulfuric acid under an argon atmosphere and the mixture was magnetically stirred in
an ice bath. About 0.5 h later, the ice bath was removed and the stirring was continued
at room temperature (23°C). ¹H-NMR was used to trace the DSs. After a determined
25 reaction time, the reaction medium was poured onto crushed ice and the resulting
precipitate was recovered by filtration, washed with deionized water until the pH
value was ~6-7.

Polymer Analysis and Measurement

Proton and carbon spectra were obtained on a Varian Unity Inova NMR spectrometer
30 operating at a proton frequency of 399.951 MHz and a carbon frequency of 100.578
MHz. Tetramethylsilane was used as the internal standard chemical shift reference.

¹H-NMR spectra of PPEK and low DS SPPEK were acquired at a temperature of

22°C in deuterated chloroform (CDCl_3). Deuterated methylene chloride (CD_2Cl_2)

was used when accurate integration values were desired for high field aromatic

signals since the signal of residual CH_2Cl_2 (5.31 ppm) did not interfere with the

aromatic region. Deuterated dimethylsulfoxide (DMSO-d_6) was the NMR solvent of

5 choice for higher DS SPPEK. Presaturation of the large water peak always present in

SPPEK at around 4 ppm improved the spectra by increasing the signal to noise ratio

and by removing interference signals from the large water absorption. ^{13}C NMR

spectra of PPEK were acquired using a 10 mm tunable broadband probe and a

concentrated sample in CDCl_3 (500mg in 3mL). A quantitative spectrum was

10 obtained using long relaxation delay and hydrogen decoupling only during the

acquisition time. IR spectra were measured on a Nicolet 520 Fourier transform

spectrometer with powder samples inside a diamond cell.

A TA Instruments thermogravimetric analyser (TGA) instrument model 2950

was used for measuring the degradation temperatures (T_d) and a TA Instruments

15 differential scanning calorimeter (DSC) model 2920 calibrated with Tin at 231.93°C

and Zinc at 419.53°C was used for measuring the glass transition temperatures (T_g).

Polymer samples for TGA analysis were preheated to 150°C at 10°C/min under

nitrogen atmosphere, held isothermally for 60 min, equilibrated at 80°C, then heated

to 800°C at 10°C/min for T_d measurement. Hence, the degradation data reported here

20 were assumed to be in the absence of moisture. Samples for DSC analysis were

initially heated at a rate of 10°C/min under nitrogen atmosphere to well below the

polymer T_d point, quenched in liquid nitrogen and then re-heated at the same rate.

Inherent viscosities were determined using an Ubbelohde viscometer for *N,N*-dimethylformamide solutions of polymer with a concentration of 0.5g/dL at 30°C.

Preparation of Membrane Films

An amount of 0.7 g sample was dissolved in 12 mL of *N,N*-dimethylformamide (DMF) or *N,N*-dimethylacetamide (DMAc) and filtered. The filtered solution was poured onto a glass plate and dried at 40°C for about two days. Residual solvent was
 5 further evaporated at 120°C under vacuum for 48 h, resulting in yellow membrane films.

Water Uptake Content Measurement and Swelling Ratio

All polymer membranes used were vacuum dried at 120°C before test. The sample films were soaked in deionized water for different time at determined temperatures.

10 Weights of dry and wet membranes were measured. The water uptake content was calculated by

$$\text{Uptake content (\%)} = \frac{\omega_{\text{wet}} - \omega_{\text{dry}}}{\omega_{\text{dry}}} \times 100\%$$

Where ω_{dry} and ω_{wet} are the masses of dried and wet samples respectively. The swelling ratio was calculated from films 7~10 cm long by:

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$$\text{Swelling ratio (\%)} = \frac{l_{\text{wet}} - l_{\text{dry}}}{l_{\text{dry}}} \times 100\%$$

Where l_{dry} and l_{wet} are the lengths of dry and wet samples respectively.

Conductivity

The proton conductivity was measured by AC impedance spectroscopy over a frequency range of 1-10⁷ Hz with oscillating voltage 50-500 mV, using a system
 20 based on a Solarton 1260 gain phase analyzer. A sample with diameter 10 mm was placed in an open, temperature controlled cell, where it was clamped between two blocking stainless steel electrodes with a permanent pressure of about 3 kg/cm².

Specimens were soaked in deionized water prior to the test. The conductivity (σ) of the samples in the transverse direction was calculated from the impedance data, using
 25 the relation $\sigma = d/RS$ where d and S are the thickness and face area of the sample respectively and R was derived from the low intersect of the high frequency semi-circle on a complex impedance plane with the Re (Z) axis.

RESULTS AND DISCUSSION

Sulfonation reaction

30 Figure 1 shows the sulfonation reaction, whereby sulfonation occurs as expected around the electron-donating ether linkage.

In general, sulfonation of polymers can be conducted with several sulfonating agents, including concentrated sulfuric acid, fuming sulfuric acid, chlorosulfonic acid and acetyl sulfate, depending on the reactivity of the polymer. For Victrex PEEK, a high DS can be achieved with concentrated sulfuric acid because of the high reactivity of the hydroquinone segment in the polymer chain. We initially attempted the sulfonation of PPEK in 95-98% concentrated sulfuric acid at both room and elevated temperatures. However, the results listed in Table 1 show that almost no sulfonation of PPEK took place at room temperature even for a very long reaction time, and only SPPEK with a very low DS was obtained at 60°C after 60 h. Sulfonation is an electrophilic reaction affected by both the electrophilicity of the sulfonating agent and the electron donating characteristics of polymer. Compared with PEEK, PPEK (as shown in Fig. 1) has more electron-withdrawing functionality per repeat unit, which decreases the reactivity of the electron-donating ether linkage of the polymer. In order to prepare SPPEK with a high DS, 27-33% fuming sulfuric acid and elevated reaction temperature were employed. As shown in Table 1, the sulfonation of PPEK in oleum at 40°C for 1 h resulted in SPPEK with a DS of 1.6. The DS didn't increase significantly over extended reaction times. This is because the sulfonic acid groups are more than 1.5 per repeating unit, reducing the electron-donating characteristic of the polymer and preventing the further introduction of more sulfonic acid groups. Although a high DS was achieved using fuming sulfuric acid, the reaction proceeded too rapidly to maintain control. Abating the reaction to room temperature reduced the DS to 1.3. However, at a DS of 1.3, SPPEK is water soluble and can not be used as a PEM in fuel cell where high humidity exists. A further reduction in temperature made the PPEK solution in fuming sulfuric acid too viscous to completely dissolve the polymer, which may result in a heterogeneous sulfonation reaction. In order to reduce the reactivity of the sulfonating agent, a mixture of 95-98% concentrated sulfuric acid and 27-33% fuming sulfuric acid was used for PPEK. As seen in Table 1, by varying the ratio of concentrated sulfuric acid to fuming sulfuric acid and the reaction time, the sulfonation reaction was more readily controlled to different DSs in the range of 0.6 to 1.23.

Table 1 Dependence of DS on the reaction conditions

Oleum/ Concentrated sulfuric acid	Reaction temperature (°C)	Reaction time (h)	DS
0/10	R.T.	95	~0
-	60	6	<0.1
4/6	R.T.	4	0.1
5/5	R.T.	1	0.6
		4	0.8
		7	0.9
6/4	R.T.	1	0.8
		4	1.0
		7	1.1
7/3	R.T.	1	1.0
		4	1.2
10/0	R.T.	1	1.3
	40	1	1.6
		2	1.6
		23	1.7

Reaction conditions: PPEK 0.5 g, mixture of sulfuric acid 10 mL, argon atmosphere. R.T.: Room Temperature

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NMR

In order to determine the sulfonation site and the DS quantitatively, the ^1H -
 10 NMR spectra of PPEK in CDCl_3 (Figure 2) and SPPEK in DMSO-d_6 (Figure 3) were characterized. The number system adopted for the protons in the polymer and the derivatives are shown in Figure 2 to 4. The careful and accurate assignments of NMR spectral signals was essential for determining both the DS and the site of sulfonation.

The starting point for peak assignment of ^1H and ^{13}C -NMR of PPEK was from the absorptions resulting from atoms surrounding the electron rich ether moiety. In a PPEK repeat unit, the hydrogen atoms at the ortho position to the ether linkage are more shielded than any other hydrogen atoms due to resonance effect of the oxygen lone pair electrons. The other functional groups present in PPEK have a deshielding effect on nearby nuclei. As a result of that, the furthest upfield signals (7.08-7.30 ppm) arise from the 4H ortho to the ether linkage. 2D COSY NMR as well as simple homonuclear decoupling experiments showed the presence of two separate spin coupling systems originating from these high field signals (Figure 2). These interactions are a result of spin-spin coupling between the hydrogen atoms at the ortho and meta positions of these phenol rings. One spin system consists of the high field ortho ether 2H absorption at 7.16-7.30 ppm coupled with the distinct signal at 7.60-7.72 ppm. The other spin system consists of the high field ortho ether 2H absorption at 7.08-7.16 ppm coupled with the multiple absorptions at 7.76-8.06 ppm.

The results of a simple ^{13}C -NMR experiment allowed formal assignment of the preceding hydrogen signals, which are key components for the accurate measurement of the DS of SPPEKs. The H-decoupled ^{13}C -NMR spectrum of PPEK is displayed in Figure 4. The carbon peaks of particular interest are the ones arising from carbon atoms directly linked to an electron-withdrawing heteroatom (N or O) causing absorptions to appear at lower field (144-195 ppm). The furthest downfield signal (193.6-194.8 ppm) is unequivocally from the ketone carbon. Further upfield, three carbon signals (155-162 ppm) arise from carbon atoms linked to electronegative oxygen atoms in the phthalazinone and ether functionalities. Finally, the less electronegative nitrogen atoms give the carbon absorptions at higher field (144-148 ppm). In one ^{13}C -NMR experiment, a spectrum was acquired with decoupling of one

of the hydrogen frequencies (7.60-7.72 ppm) responsible for the 2H at the meta position of one of the phenol rings. The resulting spectrum was compared with a ^{13}C -NMR spectrum acquired with full hydrogen coupling. Of all the downfield carbon signals, the C-O absorption at 156.0-157.1 ppm and the C-N at 147.0-148.0 ppm both lost a 6-7 Hz long range (3 bond C-C-C-H) carbon-hydrogen coupling when the hydrogen frequency 7.60-7.72 ppm had been irradiated. Hence, the long-range C-H interaction could *only* originate from C(O)-27 coupled with H-25,29 and C(N)-4 coupled with H-25,29. Decoupling of the other meta-ether hydrogen atoms H-19,23 would not affect any of the 2 C-N carbons. This experiment lead to the unequivocal assignment of H-20,22 and H-26,28 which are essential for the DS calculations. The assignment of H-25,29 and H-19,23 were also derived from this experiment. Further 1D and 2D heteronuclear (C and H) NMR experiments confirmed the previous assignments. Table 2 and 3 list the chemical shift of the ^1H and ^{13}C -NMR spectra of PPEK repectively.

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Table 2 ^1H -NMR data of PPEK in CDCl_3

Proton Number	Chemical shift δ (ppm)	Integral intensity (number of H)
H-6,7,8,9,12,13,15,19,23	7.76-8.06	9H
H-16	8.54-8.68	1H
H-20,22	7.08-7.16	2H
H-25,29	7.60-7.72	2H
H-26,28	7.16-7.30	2H

Table 3 ^{13}C -NMR data of PPEK (hydrogen decoupled) in CDCl_3 .

Carbon Number	Chemical shift δ (ppm)
C-1	158.72
C-4	147.0-148.0
C-5-10,12-16,18,19,23-25,29	125.0-137.0
C-11	144.8-146.0
C-17	193.6-194.8
C-20,22	117.6-118.6
C-21	160.2-161.4
C-26,28	119.4-120.4
C27	156.0-157.1

The hydrogen absorption at low field (8.54-8.68 ppm) is strongly believed to arise from H-16 (Figure 2) although it could not be formally proved. The chemical shift position of many peaks in ^1H and ^{13}C -NMR spectra is seen to be not averaged out as would be the case where free rotation occurs. It is believed that PPEK is hindered by restricted rotation around certain linkages. In a "restricted" structure, H-16 would have an intensity of 1, as observed, and also would be in close proximity to the phthalazinone carbonyl. The proximity to the carbonyl would deshield H-16 more than any other hydrogen atoms nearby one of the carbonyl groups of PPEK.

DS calculations from ^1H -NMR:

A ^1H -NMR of PPEK in CD_2Cl_2 was acquired and the integration value of the upfield H-26,28 and H-20,22 absorptions was set to 4.00. The integration values of the other regions of the spectra corresponded exactly (table 2) to the number of hydrogen atoms expected from the chemical structure of PPEK repeat unit.

Similarly, in a quantitative ^{13}C -NMR experiment, the upfield carbon signals C-26,28 and C-20,22 were also set to an intensity of 4.00 and the integration values of the other peaks again matched precisely. Sulfonated PPEK is only soluble in highly polar solvents and DMSO- d_6 was chosen to dissolve SPPEK. Spectra of SPPEK DS 1.03

5 and 1.63 are displayed in Figure 3. In comparison with PPEK, several new signals for the SPPEK derivatives appear with different chemical shifts due to the different solvents used and perturbation by different DSs of $-\text{SO}_3\text{H}$ groups. The ortho-ether hydrogen atoms remained at high field but their intensity decreased as they were replaced by $-\text{SO}_3\text{H}$ groups. More importantly, the ratio of low field H-16 (8.40-8.55) 10 ppm to multiple peaks 7.60-8.40 ppm (which do not include ortho-ether protons) always remained 1.00:11.00 for low and high sulfonation degrees. This is proof that substitution occurred only at the ortho-ether sites of the PPEK repeat unit. Another phenomenon supporting this is the variation in size of the small high-field signal at 7.30-7.40 ppm due to mono-substitution on SPPEK repeat units. When strong 15 electron withdrawing sulfonic acid groups are attached to benzene rings, they induce deshielding of hydrogen in the ortho and para positions. In monosubstituted SPPEK (DS~1), H-20 is deshielded by the $-\text{SO}_3\text{H}$ group hence shifted downfield. In disubstituted SPEEK repeat units (DS>1), H-20 is still deshielded by the $-\text{SO}_3\text{H}$ group present on the phenol ring but it is also shielded by the proximity through space 20 of the electron-rich oxygen atoms of the other $-\text{SO}_3\text{H}$ group nearby on the other phenol ring. The DS was simply measured by presetting the integration value of the low field hydrogen absorptions to 12H (7.50-8.60 ppm) and determining the intensity value of the upfield hydrogen signals. This value represents the number of hydrogen atoms that have not been converted into $-\text{SO}_3\text{H}$ groups and therefore subtracting it 25 from 4H (unmodified PPEK) gives a direct DS value for SPPEK.

FT-IR

Fourier Transform Infrared (FT-IR) Spectroscopy was used to confirm the pendant
5 SO_3H group on the polymer chain. Figure 5 shows the FT-IR spectra of parent PPEK
and its sulfonated derivatives with DSs of 1.03 and 1.63 respectively. In comparing
these spectra, one can see that in addition to the predicable absorptions at 3400 cm^{-1}
due to the stretching of the hydroxyls of SO_3H groups, the SPPEK absorption bands at
1020 and 1081 cm^{-1} are characteristic of the aromatic SO_3H symmetric and
10 asymmetric stretching vibrations respectively. These two characteristic peaks increase
with higher DS. In addition, the reaction can be readily followed by the signal at 1500 cm^{-1}
related to 1,4-aromatic ring substitution. Introduction of SO_3H onto the aromatic
ring induces the formation of two new adsorptions at 1471 and 1475 cm^{-1} , which at
higher sulfonation degree, completely replace the adsorption at 1500 cm^{-1} . The
15 introduction of sulfonic acid groups in the modified polymer is thus confirmed.

Thermal analysis

The thermal stabilities of the SPPEKs were determined by TGA. All the
samples were preheated at 150°C for 60 min in the TGA furnace to remove
20 moisture, then dynamic TGA experiments were run from 80 to 800°C at a heating rate
of $10^\circ\text{C}/\text{min}$ under nitrogen. Figure 5 shows the degradation curves. The parent
PPEK is a thermostable polymer of which the 5 wt% loss temperature is nearly
 500°C . For the sulfonated PPEK, there are two transitions of loss in weight. The first
one occurs at about 300°C and could be ascribed to the decomposition of the SO_3H
25 groups. In order to confirm our speculation, the weight loss from the initial point to
 455°C of SPPEK with DS of 1.0 was analyzed. A value of about 15% was obtained,
which is close to the theoretical SO_3H weight loss percentage of 16% in SPPEK with
DS of 1.0. The second thermal degradation at about 490°C is assigned to the
degradation of the main polymer chain, which is in close agreement to the weight loss

step in the TGA curve for parent PPEK. In addition, SPPEKs with higher DS lose weight more quickly than those with lower DS, in the temperature range of 300 to 460°C. It also indicates that the weight loss during this period is due to the elimination of $-\text{SO}_3\text{H}$ groups.

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The T_g of starting material is 263°C. For sulfonated product, only the T_g s of SPPEKs with DSs of 0.1 and 0.6 have been detected at 270°C and 292°C, respectively. The introduction of sulfonic groups into polymer chains leads to increased T_g s because of the increased intermolecular ionic interactions. For SPPEK with DS higher than 0.6, no T_g s have been detected because the decomposition temperatures are lower than the T_g s.

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Solubility and Viscosity

PPEK is soluble in CHCl_3 , chlorobenzene and some other chlorinated solvents, but insoluble in polar aprotic solvents such as *N,N*-dimethylformamide (DMF), *N,N*-dimethylacetamide (DMAc) and dimethylsulfoxide (DMSO). SPPEK with low DS, such as DS of 0.1 has similar solubility characteristics to the parent polymer. However, SPPEKs with high DSs are soluble in these dipolar aprotic solvents, but insoluble in chlorinated solvents. DMF was chosen as the solvent for determining the inherent viscosities $[\eta]$ of SPPEKs with high DS, which are listed in Table 4. It shows that the inherent viscosities of SPPEKs are higher compared with that of PPEK of 0.6 in chloroform and generally increase with increasing DS suggesting that the polymer chain is not degraded during sulfonation. Only SPPEK with DS of 1.63 which is obtained with 27-33% fuming sulfuric acid had a slightly reduced viscosity. A significant increase in the viscosity may arise from two factors: one is that the introduction of sulfonic acid into polymer chain caused an overall increase in the polymer molecular weight; the other is that hydrogen bonding related to sulfonic acid groups increases the molecular forces. The slight decrease in the viscosity of SPPEK with DS of 1.63 may be caused by degradation taking place during the sulfonation reaction in undiluted fuming sulfuric acid.

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Table 4 Inherent viscosities of SPPEK

DS	$[\eta]$ dL/g
0	^a 0.6
0.6	1.40
0.8	1.81
1.0	1.83
1.2	1.99
1.6	1.88

Determination conditions: SPPEK 0.5 g/dL in DMF, 30°C.

5 ^a In chloroform

Water uptake and swelling ratio

Adequate hydration of membranes is critical to fuel cell application. Water assists in the transportation of protons from the anode to the cathode. If the electrolyte membrane is too dry, its conductivity falls; on the other hand, excess water results in cathode flooding and morphological instability of membrane. The water uptake and swelling ratio of SPPEK membranes were determined by measuring the change in the mass and length before and after hydration. The results are listed in Table 5, which show that the water uptake of SPPEK increases with the DS at both room temperature and elevated temperature. At 80°C, water uptake and swelling ratio of SPPEKs with lower DS reach equilibrium values quickly. However, water uptake and swelling ratio of SPPEK with $DS \geq 1$ increase with time and some samples were swollen or dissolved in water. The molecular structure of SPPEK is composed a hydrophobic backbone and hydrophilic sulfonic acid groups. Absorbed water acts also as plasticizer, which promotes the phase separation. When the DS is are high enough, it becomes easy for the SPPEK chains to be separated by water and dissolved.

Table 5. Water uptake and swelling ratio of SPPEK

	Room temperature	80°C			
DS	3 days	2 h		24 h	
	Water uptake (%)	Water uptake (%)	Swelling ratio (%)	Water uptake (%)	Swelling ratio (%)
0.1	6	14	2.4	14	2.4
0.6	15	16	3.6	16	3.6
0.8	19	21	6.4	22	6.4
1.0	19	44	12	63	20
1.2	32	Swelling	Dissolved	—	—
1.6	100	Dissolved	—	—	—

Proton conductivity

Prior to conductivity measurements, all membrane samples were soaked in water 1 or 2 days for hydration. The effect of the DS on the conductivity of SPPEK is shown in Figure 6, which shows that the conductivity of SPPEK at room temperature increases with DS and reaches 2×10^{-2} S/cm for SPPEK with DS of 1.2. This value is similar to Nafion 117, which shows conductivity of 3×10^{-2} S/cm. Figure 7 also shows the influence of temperature on the conductivity for DS 1.0 and 1.2 SPPEKs. As can be seen, the conductivities of these two SPPEKs increase with increasing temperature up to 95°C and reach 4×10^{-2} S/cm and 6×10^{-2} S/cm, respectively. Conductivity of SPPEK with DS of 1.0 drops sharply after that, which is probably caused by the dehydration of membrane. Compared with Nafion 117, the drop in conductivity at 80°C occurs at a higher temperature for SPPEK.

CONCLUSIONS

A series of SPPEKs with different DS were prepared from PPEK with a mixture of fuming and concentrated sulfuric acid as both the solvent and sulfonating agent. The structure of SPPEK was confirmed by FT-IR and the DS of SPPEK was determined by $^1\text{H-NMR}$. As the DS of SPPEK increases, T_d decreased and T_g increased. Membrane films prepared from SPPEKs show a continuous increase in water uptake and swelling ratio with DS. Membranes prepared from SPPEK with DS of 1.0 and 1.2 show attractively high conductivity of 10^{-2} S/cm at both room temperature and elevated temperature.

Data for SPPEs and SPPEsK prepared by sulfonating PPES and SPPEsK respectively

Sulfonation Reaction of SPPEs

An amount of 2 g PPES powder was added to a mixture of 16 mL 95-98% concentrated sulfuric acid and 24 mL 27-33% fuming sulfuric acid under an argon atmosphere and the mixture was magnetically stirred in an ice bath. About 15 min later, the ice bath was removed and the stirring was continued at room temperature (23°C). $^1\text{H-NMR}$ was used to trace the DSs. After a determined reaction time, the

reaction medium was poured onto crushed ice and the resulting precipitate was recovered by filtration, washed with deionized water until the pH value was ~6-7.

Sulfonation Reaction of SPPEK

An amount of 2 g PPESK powder was added to a mixture of 16 mL 95-98% concentrated sulfuric acid and 24 mL 27-33% fuming sulfuric acid under an argon atmosphere and the mixture was magnetically stirred in an ice bath. About 15 min later, the ice bath was removed and the stirring was continued at room temperature (23°C). $^1\text{H-NMR}$ was used to trace the DSs. After a determined reaction time, the reaction medium was poured onto crushed ice and the resulting precipitate was recovered by filtration, washed with deionized water until the pH value was ~6-7.

Conductivity

Prior to conductivity measurements, all membrane samples were soaked in water 1 or 2 days for hydration. Figure 8 shows the influence of temperature on the conductivity for SPPEK and SPPEK with DS 1.0. As can be seen, the conductivity of SPPEK increases with increasing temperature up to 95°C and reaches 2×10^{-2} S/cm, then drops sharply after that, which is probably caused by the dehydration of membrane. The conductivity of SPPEK increases with increasing temperature up to 106°C and reaches 4×10^{-2} S/cm, then drops. Compared with Nafion 117, the drop in conductivity at 80°C occurs at a higher temperature for SPPEK.

Proton conductivity

Prior to conductivity measurements, all membrane samples were soaked in water 1 or 2 days for hydration. The effect of the DS on the conductivity of SPPEK is shown in Figure 9, which shows that the conductivity of SPPEK at room temperature increases with DS and reaches 2×10^{-2} S/cm for SPPEK with DS of 1.2. This value is similar to Nafion 117, which shows conductivity of 3×10^{-2} S/cm. Figure 9 also shows the influence of temperature on the conductivity for DS 1.0 and 1.2 SPPEKs. As can be seen, the conductivities of these two SPPEKs increase with increasing temperature up to 95°C and reach 4×10^{-2} S/cm and 6×10^{-2} S/cm, respectively. Conductivity of SPPEK with DS of 1.0 drops sharply after that, which is probably caused by the dehydration of membrane. Compared with Nafion 117, the drop in conductivity at 80°C occurs at a higher temperature for SPPEK.

Data for SPPEK prepared by polymerization of sulfonated monomers

(This demonstrates that there are alternative methods of producing poly(phthalazinones)

The random sulfonated poly(phthalazinone ethers) have also been obtained via direct polymerization reactions with the sulfonation sites on the deactivated aromatic rings as depicted in scheme 1.

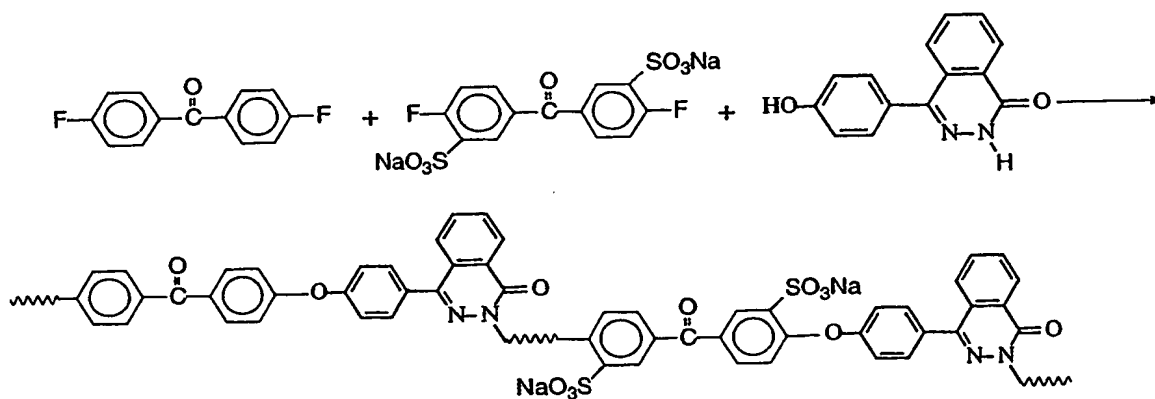
Polymerization Reaction

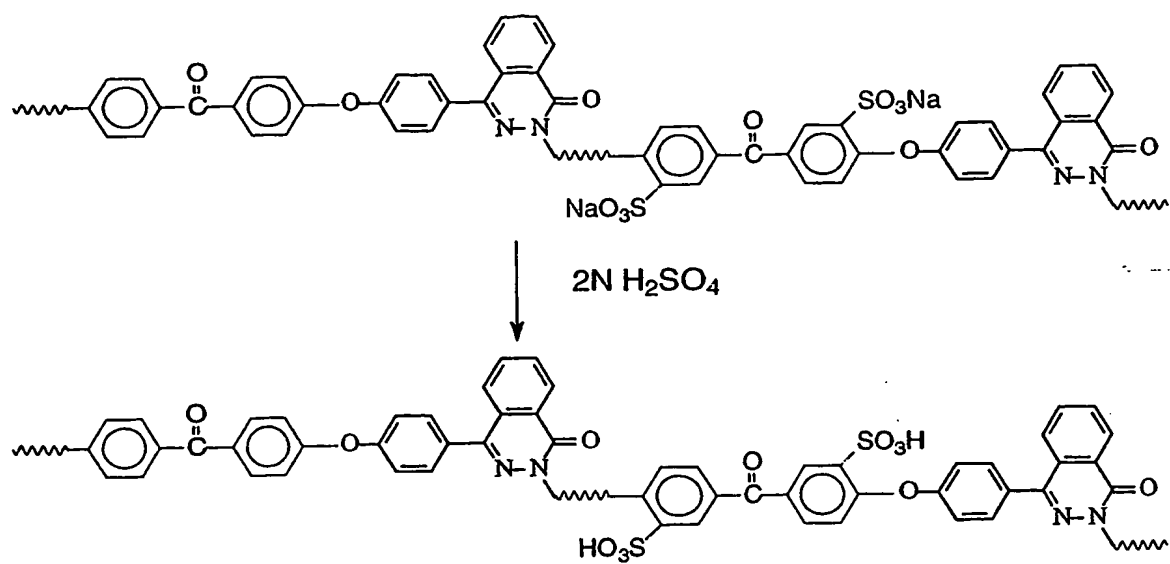
The SPPEKs synthesized with various compositions are denoted as SPPEK-*n*, where *n* means the percentage content of SDFK in feed of DFK and SDFK. The synthesis of SPPEK-50 is used as a typical example. To a three-neck flask with a magnetic stirrer, a Dean-Stark trap and condenser, and an Argon inlet, 1.102 g DFK (5.05 mmol), 2.133 g SDFK-Na (5.05 mmol), 2.383 g DHPZ (10 mmol), and 1.8 g potassium carbonate (13 mmol) were added. Then 18 mL of NMP and 25 mL chlorobenzene were charged into the reaction flask under an argon atmosphere. The reaction mixture was heated to 140°C. Upon dehydration and removal of chlorobenzene, the reaction temperature was increased to 170-175°C. After a period of 5-7 h, when the solution viscosity had obviously increased, several milliliters of NMP was added to dilute the solution and kept it for further 3-5 h. Then, the mixture was cooled to 100°C and coagulated in ethanol, water or acetone. After recovering and drying the product, SPPEKs were purified by dialysis for 7 days, using a membrane-cellulose dialysis tube (SPECTRUM) with a molecular weight cut off value of 3500.

Preparation of Membrane Films

An amount of 1 g SPPEK in sodium form was dissolved in 12 mL of *N,N*-dimethylacetamide (DMAc) and filtered. The filtered solution was poured onto a glass plate and dried at 40°C for about two days. Residue solvent was further evaporated at 120°C under vacuum for 48 h. resulting in yellow membrane films. The acid form membrane films were obtained by doping the sodium form membrane films into 2 N H₂SO₄ for 48 h, followed by doping in deionized water for 48 h, and vacuum drying at 100°C for 24 h.

Scheme 1 Synthetic formula for SPPEK



Scheme 2 Preparing H-form membranes

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